Environment Pollution and Climate Change

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Impact of Chlorofluorocarbons on Environment and Climate Change

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Abstract

similar to that of simpler alkanes.

Keywords: Hydrochloro uorocarbons; Chloro uorocarbons

Introduction

e methane-derived CFCs deviate from perfect tetrahedral symmetry because the uorine and chlorine atoms di er greatly in size and e ective charge from hydrogen and from one another. Changing the number and identity of the halogen atoms can alter the physical properties of CFCs and HCFCs. ey are volatile in general, but less so than their parent alkanes. e halides' molecular polarity, which causes intermolecular interactions, is responsible for the decreased volatility. As a result, while uoromethanes boil between 51.7 (CF2H2) and 128°C (CF4), methane boils at 161°C. Because chloride is even

umberafCcloridecorrelated.Duetotheirlowtoxicity,reactivity,and ammability,CFCsandHCFCsareutilizedinnumerousapplications. emajorityofpermuta monomer that is transformed into Te on, billions of kilograms of chlorodi uoromethane are produced annually.

CFCs' e ects on the atmosphere go beyond their role as ozone-depleting chemicals. e earth's atmosphere is kept warm at that wavelength by infrared absorption bands. CFCs and other unreactive uorine-containing gases like per uorocarbons, HFCs, HCFCs, bromo uorocarbons, SF6, and NF3 create a "super" greenhouse e ect because of the strength of CFC aubeyond themeyony, 35ecareieieiene,

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of the ozone layer over Antarctica. By the end of the century, 12 nations of the European Community had agreed to prohibit the production of all CFCs. When diplomats met in London in 1990, they decided to call for the complete elimination of CFCs by the year 2000, which would signi cantly strengthen the Montreal Protocol. CFCs ought to have been completely eliminated from developing nations as well by 2010.

Prices for CFCs have skyrocketed as a result of the fact that recycling